

Titration Lab Procedures

1. Rinse out your 2 burets and your 250 mL Erlenmeyer flask with DI water. Retrieve your acid and base beakers from the front of the room and prepare your burets as your teacher has instructed. (Do not overfill your burets)
2. Record your **initial volume** readings on the burets of your acid and your base. You will eventually need to record the final readings for your acid and your base, but they may change throughout the experiment.
3. Fill your E-flask with approximately **20mL of ACID** (Record the exact volume added). Add 2-3 drops of PPT to the flask as well.
4. Your task will be to ***slowly*** (because the M_b is unknown) **add base** to your flask until you reach your end point (when the indicator has just started to turn pink, the solution will be very faintly pink). Don't forget to swirl your flask and give a few seconds for a reaction to occur.
5. If you add too much base & the solution is bright pink, you must go back and add more acid (NOTE: this will change your volume of your acid!)
6. Record the final volumes of your amount of acid and base added.

Post Lab Questions

1. How did you know to end the titration? What is this point called?
2. What point do we assume is the same as the end point of a titration?
3. Using your known volumes & molarity, calculate the molarity of the base.
4. Why are we allowed to use the equation you used in the previous question in this scenario?
5. What was your percent error for this lab and what are some reasons for error?